A cobalt(II) iminoiodane complex and its scandium adduct: mechanistic promiscuity in hydrogen atom abstraction reactions†

Subrata Kundu, a Petko Chernev, b Xenia Engelmann, a Chan Siu Chung, c Holger Dau, b Eckhard Bill, d Jason England, xc Wonwoo Nam,xe and Kallol Raya

In addition to oxometal \([M^{m=O}]\) and imidometal \([M^{m=NR}]\) units, transient metal–iodosylarene \([M^{m=2\text{–}O=\text{IPh}}]\) and metal–iminoiodane \([M^{m=2\text{–}O=\text{IPh}}]\) adducts are often invoked as a possible "second oxidant" responsible for the oxo and imido group transfer reactivity. Although a few metal–iodosylarene adducts have been recently isolated and/or spectroscopically characterized, metal–iminoiodane adducts have remained elusive. Herein, we provide UV-Vis, EPR, NMR, XAS and DFT evidence supporting the formation of a metal–iminoiodane complex 2 and its scandium adduct 2-Sc. 2 and 2-Sc are reactive toward substrates in the hydrogen-atom and nitrogen transfer reactions, which confirm their potential as active oxidants in metal-catalyzed oxidative transformations. Oxidation of para-substituted 2,6-di-tert-butylphenols by 2 and 2-Sc can occur by both coupled and uncoupled proton and electron transfer mechanisms; the exact mechanism depends on the nature of the para substituent.

Iodosobenzene \((\text{PhI}=\text{O})\) and iminoiodanes \((\text{PhI}=\text{NR})\) are an important class of group transfer reagents in organic synthesis, and they are often used in conjunction with transition metal-based catalysts.1 High-valent oxometal \([M^{m=O}]\) and imidometal \([M^{m=NR}]\) units are generally accepted to be key reactive intermediates in these reactions.2 However, transient metal–iodosylarene \([M^{n=2\text{–}O=\text{IPh}}]\) and metal–iminoiodane \([M^{n=2\text{–}O=\text{IPh}}]\) adducts have also been suggested as a possible "second oxidant" responsible for this oxo and imido group transfer reactivity. Nam et al. reported the first isolation and spectroscopic characterization of an iodosobenzene adduct, namely \([\text{porph}]\text{Fe}^{\text{III}}(\text{OIPh})\]^+, from the reaction of \([\text{porph}]\text{Fe}^{\text{IV}}\) with iodosobenzene.3c Very recently, McKenzie et al. provided the molecular structure of a non-heme iron(ii) iodosylarene adduct by X-ray crystallography.4d Fujii et al.5 also reported structural and spectroscopic evidence in support of bis(iodosylarene) coordination to a manganese(rv)-salen center. Notably, the aforementioned isolated \([M^{n=2\text{–}O=\text{IPh}}]\) adducts were all found to be reactive toward substrates in both hydrogen- and oxygen-atom transfer reactions, which confirms their potential as active oxidants in metal-catalyzed oxidative transformations.3 In contrast, isoelectronic \([M^{n=2\text{–}N=\text{IPh}}]\) complexes have remained elusive, and their presence as transient reactive intermediates has been forwarded on the basis of labelling experiments alone.4a

Herein, we describe the spectroscopic characterization of a novel \([\text{TMG}_3\text{tren}]\text{Co}^{\text{III}}(\text{PhInts})]^+ (2) complex \([\text{PhInts} = \{\text{N}(\text{p-toluenesulfonyl})\text{imino}\}[2\text{-tert-butylsulfonyl}]\text{phenylidodinate}\);\textcolor{red}{5\text{TMG}}_3\text{tren} = \text{tris}[2\text{-N-tetramethylguanidyl}][\text{ethyl}]\text{amine}\) and its scandium adduct \([\text{TMG}_3\text{tren}]\text{Co}^{\text{III}}(\text{PhInts}(\text{Sc}(\text{OTf})_3))^+ (2\text{-Sc}, \text{OTf} = \text{CF}_3\text{SO}_3\text{)}\), which are able to oxidize the C–H and O–H bonds of a variety of substrates. Additionally, we demonstrate that the presence of Sc\(^{3+}\) in 2-Sc initiates a two-step reaction mechanism for the oxidation of benzyl alcohol (PhCH_2OH), involving a [Co(III)(PhInts)⋯Sc\(^{3+}\)⋯O(H)⋯CH_2Ph] association event prior to C–H bond cleavage. In the case of para-substituted phenols two different mechanistic pathways have been established for reactions with 2 and 2-Sc: a concerted proton-coupled electron transfer (PCET), and a proton transfer followed by electron transfer (PT-ET). The nature of the phenol para-substituent is shown to control the mechanism of oxidation.

A combination of \([\text{TMG}_3\text{tren}]\text{Co}^{\text{III}}(\text{OTf})\]^– (1) with \(^{8}\text{PhInts}\) in acetone at −40 °C led after 6000 seconds to the formation of an orange complex 2 with absorption maxima \(\lambda_{\text{max}} (\varepsilon_{\text{max}})\) centered at 420 nm [1150 M\(^{−1}\) cm\(^{−1}\)] and 820 nm [130 M\(^{−1}\) cm\(^{−1}\)] (Fig. S1†). The absorption spectrum of 2 is clearly distinct from 1 + \(^{8}\text{PhInts}\) (Fig. 1), and titration experiments (Fig. 1

1| View Issue
2| View Journal
3| View Article Online
4| This journal is © The Royal Society of Chemistry 2016
5| Accepted 11th July 2016
6| Received 9th May 2016
7| DOI: 10.1039/c6dt01815g
8| Cite this: Dalton Trans., 2016, 45, 14538
10| Creative Commons Attribution 3.0 Unported Licence.

In Fig. 1, the electronic spectra of 1 (solid trace), 4PhINTs (dash-dotted trace), 2 (dashed trace), and 2-Sc (dotted trace) in acetone at -40 °C. Inset: Plot of the absorbance of the 420 nm band of 2 against the equivalents of 4PhINTs added to 1. Notably, 2 and 2-Sc can be generated in CH2Cl2 with Imax values identical to that observed in acetone.

Fig. 2 Fourier-transformed Co K-edge EXAFS spectra of 2 (A) and 2-Sc (B) (experimental data: dotted line, simulation: solid line). Insets show the corresponding EXAFS data in a wave-vector scale before the Fourier transformations.
mass spectrum (ESI-MS; Fig. S9†) of both 2 and 2-Sc exhibits a prominent peak at m/z = 496.15, with a mass and isotope distribution pattern corresponding to \([\text{TMG}_3\text{tren})\text{Co}^{II}-(\text{sPhINTs})^2\)²¹.

DFT calculations were also performed in order to obtain further insights into the electronic and geometric structures of 2. The binding of sPhINTs to the cobalt center in 2 can in principle occur either via the sulfonamide group (O-bound or N-bound) or alternatively via one of the O-atoms of the sulfone group (Scheme 1 and Table S4†). The most stable structure of the experimentally observed spin-quartet ground (S = 3/2) state was calculated to contain the O-bound sulfonamide binding mode of sPhINTs (Scheme 1, inset), with the respective energies (Table S5†) of the geometry optimized structures of the more sterically disfavoured N-bound sulfonamide and O-bound sulfone isomers being higher by 15.3 and 26.9 kcal mol⁻¹ respectively. Consistent with this conclusion, the DFT calculated average Co–N/O distance of 2.10 Å (corresponding to four N-donors of TMG₃tren and one O-donor of sPhINTs) in an S = 3/2 O-bound sulfonamide \([\text{TMG}_3\text{tren})\text{Co}^{II}-(\text{sPhINTs})^2\)²¹ complex (Tables S1 and S4†) best matches the EXAFS determined Co–N/O distance of 2.05 Å for 2. The calculated average Co–N distances for the corresponding N-bound sulfonamide and O-bound sulfone structures are significantly longer at 2.17 Å and 2.14 Å, respectively.

We then compared the reactivities of 2 and 2-Sc in hydrogen atom transfer (HAT) reactions from the hydrocarbon substrates xanthene, 9,10-dihydroanthracene (DHA), 1,4-cyclohexadiene (CHD), benzyl alcohol (PhCH₂OH) and 1-benzyl-1,4-dihydronicotinamide (BNAH). In all cases, addition of a substrate led to significantly accelerated decay of the chromophore associated with 2 and 2-Sc. In the presence of an excess substrate this proceeded with first-order kinetics and the resulting effective rate constants \(k_{eff}\) were found to be linearly dependent upon substrate concentration, thereby yielding second-order rate constants. Such reaction kinetics are consistent with the direct reaction of 2 and 2-Sc with a substrate or alternatively a rapid pre-equilibrium yielding an alternative active oxidant. Intermediacy of a more traditional imidometal \([\text{TMG}_3\text{tren})\text{Co}^{IV}-(\text{NTs})^2\) oxidant would, therefore, require rapid reversible N–I bond scission. Given that there is no precedence for this and the high energy calculated for the N-bound isomer, which would be a necessary intermediate in its formation, this seems highly improbable. As a consequence, we postulate that the O-bound sulphonamide isomer is the active oxidant.

Table 1  Substrate reactivity studies for 2 and 2-Sc

<table>
<thead>
<tr>
<th>Substrate</th>
<th>BDE₇-H (kcal mol⁻¹)</th>
<th>Product (% yield)⁶</th>
<th>(k_2) (10⁻² M⁻¹ s⁻¹) at −40 °C</th>
<th>(k_2) (10⁻² M⁻¹ s⁻¹) at −20 °C</th>
</tr>
</thead>
<tbody>
<tr>
<td>BNAH</td>
<td>67.6</td>
<td>—</td>
<td>210.1</td>
<td>190.5</td>
</tr>
<tr>
<td>Xanthene</td>
<td>74.0</td>
<td>—</td>
<td>2.06</td>
<td>1.14</td>
</tr>
<tr>
<td>CHD</td>
<td>76.0</td>
<td>Benzene (65)</td>
<td>1.32</td>
<td>1.38</td>
</tr>
<tr>
<td>DHA</td>
<td>76.3</td>
<td>Anthracene (60)</td>
<td>0.74</td>
<td>1.17</td>
</tr>
<tr>
<td>PhCH₂OH</td>
<td>80.0</td>
<td>PhCHO (52)</td>
<td>0.20</td>
<td>0.37⁷⁸</td>
</tr>
<tr>
<td>Ph₃P</td>
<td>—</td>
<td>Ph₃P=NTs (100)</td>
<td>53.1</td>
<td>3.21</td>
</tr>
<tr>
<td>X-DTBP X =</td>
<td>BDE₁-O (kcal mol⁻¹)</td>
<td>σ</td>
<td>(pK_a)</td>
<td>(pK_a)</td>
</tr>
<tr>
<td>MeO</td>
<td>78.3</td>
<td>−0.78</td>
<td>14.82</td>
<td>100.8</td>
</tr>
<tr>
<td>Me</td>
<td>81</td>
<td>−0.31</td>
<td>14.77</td>
<td>2.42</td>
</tr>
<tr>
<td>t-Bu</td>
<td>81.2</td>
<td>−0.26</td>
<td>14.82</td>
<td>1.97</td>
</tr>
<tr>
<td>H</td>
<td>82.1</td>
<td>0</td>
<td>14.22</td>
<td>0.43</td>
</tr>
<tr>
<td>CHO</td>
<td>83.1</td>
<td>0.40</td>
<td>9.36</td>
<td>40.12</td>
</tr>
<tr>
<td>MeC(O)</td>
<td>83.1</td>
<td>0.48</td>
<td>10.27</td>
<td>31.29</td>
</tr>
<tr>
<td>COOH</td>
<td>83.4</td>
<td>0.55</td>
<td>—</td>
<td>63.47</td>
</tr>
<tr>
<td>CN</td>
<td>84.2</td>
<td>0.66</td>
<td>10.15</td>
<td>35.82</td>
</tr>
</tbody>
</table>

⁶Yields are for complex 2 and calculated based on the amount of the starting complex 1; yields for reactions with 2-Sc are comparable to 2.

⁷The rate was determined using less than 150 equiv. of benzyl alcohol.
Product analysis of the reaction mixtures of DHA, CHD and PhCH₂OH reactions revealed the formation of the corresponding dehydrogenated products (Table 1) in 52–65% yields; p-toluenesulfonylamine (ArSO₂NH₂; see Fig. 3 and 4) is also detected as a reaction product (80–85% yield) in each case. Control reactions of the substrate with sPhINTs alone or with sPhINTs + Sc³⁺ did not lead to any oxidation products. Thus, sPhINTs is activated upon binding to the [(TMG₃tren)Co⁴⁺] core. Interestingly, the second order rate constants (k₂) determined at −40 °C are comparable both in the absence and presence of Sc³⁺ ions (Table 1). This is in contrast to the previous report of the enhancement of the oxidizing capabilities of the corresponding metal-oxo intermediates⁶,⁷ in the presence of various Lewis-acids. Furthermore, the k₂ values form a linear correlation (Fig. S10†) with the C–H bond dissociation energies (BDEc–H)⁵ of the substrates, which indicates that C–H bond homolysis is the rate determining step (r.d.s) in these reactions.

In order to obtain additional mechanistic insights, deuterium kinetic isotope effects (KIE) on the second-order rate constant k₂(C–H)/k₂(C–D) were measured for reactions of xanthene and DHA with 2. Values of 1.87 and 1.66 were determined at −40 °C for xanthene and DHA (Fig. S11A and B†), respectively, indicating that proton transfer is involved in the rate determining step. These KIE values are significantly smaller than those typically seen for high-valent metal-oxo and metal-imido mediated HAT reactions (3–25).¹,² This is again consistent with the [[(TMG₃tren)Co⁴⁺−][sPhINTs]]²⁺ core in 2 not evolving to a reactive high-valent [[(TMG₃tren)Co⁴⁺−(NTs)]²⁺ species prior to the reaction, and supports the suggestion that it is able to act directly as the reactive species in these oxygenation reactions.

Although the presence of Sc³⁺ did not significantly perturb the rate of C–H bond homolysis mediated by the [[(TMG₃tren)Co⁴⁺−][sPhINTs]]²⁺ core, it does affect a change in the mechanism in the reaction with PhCH₂OH. As expected, the addition of large excess of PhCH₂OH (30–1000 equivalents) to pre-formed solutions of 2 at −40 °C caused pseudo-first order decay (Fig. S12†) of the absorption bands associated with this complex at 420 and 820 nm. Furthermore, the rate of decay was found to increase linearly with increasing PhCH₂OH concentration (Fig. 3A), thereby affording a second order rate constant (k₃) of 0.0020 M⁻¹ s⁻¹. The pseudo-first order rate constants (k₃obs) for the reaction of 2–Sc with 30–150 equivalents of PhCH₂OH were also observed to follow a linear correlation, and a second order rate constant (k₃) of 0.0037 M⁻¹ s⁻¹ at −40 °C was measured (Table 1; Fig. 3B inset). However, at higher concentrations (>150 equivalents) the correlation became non-linear with saturation behavior being obtained (Fig. 3B). The simplest model to account for these findings is to invoke the presence of a relatively fast equilibrium preceding the rate-determining step.¹⁰ More specifically, PhCH₂OH reversibly binds to 2–Sc (Fig. 3) to yield a substrate bound intermediate [[(TMG₃tren)Co⁴⁺−][sPhINTs]]…Sc³⁺…−O[H]−CH₄Ph, which then undergoes C–H bond cleavage by a rate determining H-atom abstraction process. At high PhCH₂OH concentrations (>150 equivalents), the equilibrium shifts completely towards the substrate-bound species and the rate becomes independent of PhCH₂OH concentration, thereby explaining the saturation kinetics.
This notion is further reinforced by the absence of saturation kinetics for the other substrates, which do not possess a ligating atom, and 2, which lacks a Lewis acidic binding site.

Mechanistic studies for the reaction of 2 and 2-Sc with para-substituted 2,6-di-tert-butylphenol ([X-DTBP]: X = –OMe, –Me, –Bu, –H, –CN, –CHO, –COOH, –MeC(=O)]) reveal that the second-order rate constants depend markedly on the electron-donating/-withdrawing properties of the para-substituents (Table 1). For all the substituted phenols studied, second-order kinetic behavior (Fig. S13†) and identical reaction products, namely ArSO2NH2 and two equivalents of the corresponding phenoxal radical, were observed. However, when the rates obtained for both 2 and 2-Sc are plotted as a function of σ of the para-substituents a linear correlation with a negative Hammett slope (Fig. S14†), as previously reported for the reaction of metal-oxo- species and superoxo intermediates with phenol substrates, is obtained only for the electron-donating substituents (X = –OMe, –Bu, –Me, and –H). The k2 values for the electron withdrawing substituents [X = –CN, –CHO, –COOH, –MeC(=O), –Br] appear significantly above the trend line, thereby hinting at a change in the mechanism for these substrates.

In response, we also plotted the rate constants against the O–H bond dissociation energy (BDE). Once again, a linear correlation with a negative slope comparable to that obtained for the metal-oxo- and -dioxyn intermediates was observed for the phenols with electron donating substituents, but the data for the phenols with electron withdrawing groups appeared outside the (Fig. 4A and S15A†). Plotting log k2 versus pKa (Fig. 4B and S15B†) revealed that for X-DTBP with X = –CHO, –C(O)Me, –CN, –Br, and –H, the rate (log k2) decreased linearly with decreasing acidity (increasing pKa), whereas the rates for the electron donating substituents –MeO, –Me, and –Bu scatter irregularly. From the above studies we can conclude that 2 and 2-Sc oxidize acid phenols (X-DTBP; X = –CN, –CHO, –COOH, –C(O)Me) via a stepwise proton transfer and electron transfer mechanism, with the proton transfer (PT) being effectively rate determining. This is likely to involve an initial proton transfer to solvent (acetone) to give a phenolate anion, which may bind to the Co(u) center and is much easier to oxidize than the phenol itself. The resulting rate acceleration, therefore, parallels the extent of ionization (acyd) of the phenol. In contrast, less acidic phenols like MeO-DTBP, Me-DTBP, and 4-Bu-DTBP, proceed via a concerted proton-coupled electron transfer (PCET) mechanism. Interestingly, data for 2,6-di-tert-butylphenol (H-DTBP) falls in all the plots of the second order rate constant (log k2) versus the thermochemical parameters BDEO−, pH and pKa (Fig. 4A, B and S15A, B†), thereby locating H-DTBP on the mechanistic borderline between the concerted PCET and stepwise PT-ET reaction pathways.

The reaction of excess triphenylphosphine (30–60 equivalents) with preformed 2 and 2-Sc at ~40 °C leads to the quantitative formation of triphenylphosphonium tosylazaylde (Ph3P−=NTs) and regeneration of the starting CoII complex 1. Interestingly, the second order rate constants (k2) determined for the reactions of 2 and 2-Sc with PPh3 are 0.531 and 0.0321 M−1 s−1, respectively, demonstrating that the group-transfer reactivity of 2 is decelerated by factors of ~15 upon binding of Sc3+ ions (Table 1; Fig. S16 and S17†).

**Conclusions**

In conclusion, we have provided spectroscopic evidence supporting formation of a metal-iminoiodane complex 2 and its scandium adduct 2-Sc. A novel binding mode of PhINTs via the sulfonamide O-atom is predicted in 2 based on DFT calculations. Both 2 and 2-Sc possess cobalt in a +2 oxidation state. This is in sharp contrast to the analogous oxo chemistry, where the presence of scandium led to the stabilization of an unusual {CoIV–O–Sc3+} core. The formation of the isoelectronic {CoIV–N(Ts)–Sc3+} core in 2-Sc, which would require a short Co–N(Ts) distance, is presumably prevented by the additional steric demands of the tosyl group, which is absent in the oxo chemistry. The reaction of 2 and 2-Sc with 4-substituted 2,6-di-tert-butylphenol is shown to proceed via a stepwise PT-ET mechanism for electron withdrawing substrates. Although PT-ET mechanisms have previously been invoked for the oxidation of phenols by organic radicals, this is the first experimental evidence for a metal complex mediated variant. Additionally, the presence of Sc3+ in 2-Sc is shown to promote the formation of a precursor complex during the oxidation of benzyl alcohol, thereby demonstrating the cooperativity of two metal centers in promoting substrate oxidation reactions. Taken together, the present study expands our understanding of metal-mediated oxidation reactions using iminoiodanes, with metal-iminoiodane adducts being demonstrated to be a second plausible oxidant, in addition to the often invoked high-valent metal-imido reactive intermediates in hydrogen atom abstraction and group transfer reactions.

We gratefully acknowledge financial support of this work from the Cluster of Excellence “Unifying Concepts in Catalysis” (EXC 314/2), Berlin. K. R. also thanks the Heisenberg-Programm of the Deutsche Forschungsgemeinschaft for financial support. W. N. acknowledges financial support from the NRF of Korea through the CRI (NRF-2012R1A3A2048842) and GRL (NRF-2010-00353). JE is thankful to the NAP fellowship of the Nanyang Technological University.

**Notes and references**


---

**Communication**

Dalton Transactions

This article is licensed under a Creative Commons Attribution 3.0 Unported Licence.


7 Owing to the structural complexity of 2-Sc, the actual binding site of scandium is not clear at present.


